



Corrigendum

Corrigendum to “Comparative adsorption of ciprofloxacin on sugarcane bagasse from Ecuador and on commercial powdered activated carbon” [Sci. Total Environ., volume 750 (2021) Start page 477–End page 488/141498]



María E. Peñafiel ^{a,*}, José M. Matesanz ^b, Eulalia Vanegas ^a, Daniel Bermejo ^a, Rosa Mosteo ^b, María P. Ormad ^b

^a Center for Environmental Studies, Department of Applied Chemistry and Production Systems, Faculty of Chemical Sciences, University of Cuenca, Cuenca, Ecuador

^b Water and Environmental Health Research Group, Environmental Sciences Institute (IUCA), Department of Chemical Engineering and Environmental Technology, School of Engineering and Architecture (EINA), University of Zaragoza, Zaragoza, Spain

The authors regret that the printed version of the above article contained a number of errors. The correct and final version follows. The authors would like to apologise for any inconvenience caused.

Table 1

List of equilibrium and kinetic models.

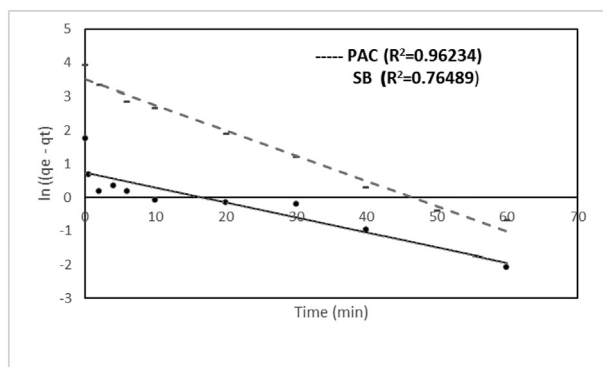
	Model	Equation	Description	Reference
Equilibrium models	Langmuir	$\frac{C_e}{q_e} = \frac{C_e}{q_m} + \frac{1}{K_L q_m}$ $R_L = \frac{1}{1 + K_L C_e}$	($R_L = 0$), favorable ($0 < R_L < 1$), linear ($R_L = 1$) or unfavorable ($R_L > 1$)	(Food and Hammed, 2010; Xiao et al., 2021)
	Freundlich	$\log q_e = \log K_F + \frac{1}{n} \log C_e$		(Food and Hammed, 2010)
	Brunauer-Emmet-Teller (BET)	$q_e = \frac{q_s C_{BET} C_e}{(C_e - C_s) [1 + (C_{BET} - 1) (\frac{C_e}{C_s})]}$		(Ebadi et al., 2009)
Kinetic models	Pseudo first order	$\ln(q_e - q_t) = \ln q_e - k_1 t$		(Yuh-Shan, 2004)
	Pseudo second order	$\frac{t}{q_t} = \frac{1}{K_2 q_e^2} + \frac{t}{q_e}$		(Ho and McKay, 1999)
	Weber and Morris	$q_t = K_1 t^{0.5} + C$ $R_i = 1 - \left(\frac{C}{q_{ref}}\right)$	R_i is the ratio of the initial adsorption amount (C) to the final adsorption amount (q_{ref}) ^a	(Wu et al., 2009)

^a q_{ref} (mg g^{-1}) is the solid phase concentration at time $t = t_{ref}$ for an adsorption system; t_{ref} is the longest time in the adsorption process.

Table 5

Kinetic parameters of CPX adsorption on SB and PAC.

Adsorbent	Experimental q_e (mg g^{-1})	Pseudo-first order			Pseudo-second order			Weber and Morris	
		q_e (mg g^{-1})	K_1 (min^{-1})	R^2	q_e (mg g^{-1})	K_2 ($\text{g mg}^{-1} \text{min}^{-1}$)	R^2	C (mg g^{-1})	R^2
SB	5.72	5.73	0.046	0.76	5.81	0.090	0.99	2.87	0.92
PAC	50.12	50.0	0.083	0.96	51.2	0.0045	0.99	27.31	0.91

Fig. S2. Pseudo-first-order kinetic plots for the adsorption of CPX on SB and PAC (CPX: 20 mg L⁻¹, 30 °C, PAC: 0.3 g L⁻¹; SB: 3 g L⁻¹).

Declaration of competing interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.